



**COMPOSITION AND METHOD FOR THE IN SITU
REMOVAL OF SCALE FROM A SUBSTRATE**

FIELD OF THE INVENTION

- 5 [0001] The subject invention generally relates to a composition and method for the in situ removal of scale from a substrate. More specifically, the subject invention relates to a composition and method for the in situ removal of silicate-containing scale from interior surfaces of a boiler.

BACKGROUND OF THE INVENTION

- 10 [0002] Compositions and methods for removing scale are known in the art. It is understood that steam is produced for various residential, commercial, and industrial purposes. For instance, steam is produced in a heat exchange apparatus, such as a steam boiler, to provide heat for residential, commercial, and industrial buildings. Steam is also
15 produced to generate power for various residential, commercial, and industrial applications. In the production of steam, large quantities of water are heated to temperatures that exceed the boiling point of water. As such, minerals and salts, which are dissolved in the water, are deposited on interior surfaces of the steam boiler. These deposits are known in the art as scale.

- 20 [0003] Scale is undesirable as it decreases the thermal efficiency of the steam boiler. Scale may also possess corrosive properties that can degrade the interior surfaces of the steam boiler and shorten the useful operating life of the steam boiler. It is generally understood that there are three classifications of scale. First, there is scale, typically an iron oxide, that is generated as the steam boiler corrodes. Second, there is scale that is

a phosphate, carbonate, or sulfate of either calcium or magnesium. This second classification of scale is most common. Third, there is scale that is a silicate-sulfate complex of calcium, magnesium, aluminum, and other various metals. This third classification of scale is typically the most difficult scale to remove from the interior surfaces of the steam boiler.

[0004] The conventional compositions and methods of the prior art are primarily directed at removal of the first two classifications of scale. For example, conventional compositions include acids targeted to remove phosphate, carbonate, and sulfate scale of calcium and/or magnesium. It is known throughout the art that these acid compositions are ineffective in removing the third classification of scale from the steam boilers. More specifically, these acid compositions are ineffective in removing the silicate-sulfate complexes of calcium, magnesium, aluminum, and other metals. Other conventional compositions rely exclusively on chelants, such as ethylenediaminetetraacetic acid (EDTA) and nitrilotriacetic acid (NTA), to remove the phosphate, carbonate, and sulfate-based scale. As with the acid compositions, these chelant only-based compositions are ineffective in removing the third classification of scale and even require several days to remove the other classifications of scale. Other conventional compositions rely on citric acid for the removal of scale. More specifically, these conventional compositions rely on citric acid in an overall acidic environment, i.e., at pHs < 7, to remove only the first classification of scale, the iron oxide scale.

[0005] Due to the overall ineffectiveness of the above compositions in removing the silicate-sulfate complexes of calcium, magnesium, aluminum, and other metals, the many industries that utilize steam for heat, power, and other reasons, have had to resort to impractical, time-consuming, and expensive methods to prevent the deposition of this

third classification of scale. These industries have even had to resort to physical removal of this scale. For example, some in industries must continuously treat their water to remove the minerals and salts that form the scale from the water, or to prevent the deposition of the scale. Other industries take the steam boiler out of service for certain periods of time and employ personnel to physically remove the scale while the steam boiler is out of service.

[0006] In sum, the prior art compositions described above are characterized by one or more inadequacy. Due to the inadequacies identified in the prior art, it is desirable to provide a novel composition and method that use simple ingredients, such as citric acid in an overall basic environment established by an alkali metal hydroxides, to effectively remove silicate-containing scale from the interior surfaces of a boiler.

SUMMARY OF THE INVENTION

[0007] A composition and method for the in situ removal of scale from a substrate are disclosed. Specifically, the composition and method of the subject invention are effective in removing silicate-containing scale, formed as silicate-sulfate complexes of calcium, magnesium, aluminum, and other metals, from the interior surfaces of boilers. The substrate is exposed to the composition to remove the scale. The composition includes a chelating agent (A) and a basic agent (B). The chelating agent (A) has at least two carboxylic acid functional groups and is preferably citric acid. The basic agent (B) is an alkali metal hydroxide. More specifically, the basic agent (B) is preferably potassium hydroxide, sodium hydroxide, or a combination of the two. The chelating agent (A) is present in the composition in an overall basic environment having a pH of from 7 to 14 established by the amount of the alkali metal hydroxide.

[0008] The method of the subject invention circulates the composition to contact the substrate. The method also includes other steps, such as heating the composition, filtering the composition, and, if necessary, re-circulating the composition, to ensure sufficient removal of the silicate-containing scale from the substrate. As the composition circulates to contact the substrate, the chelating agent (A) interacts with and extracts atoms of calcium, magnesium, aluminum, and other metals from the silicate-containing scale, and the overall basic environment in the composition enables these metal atoms to precipitate out of the composition as solid particulates.

[0009] Accordingly, using simple chemical ingredients, the subject invention provides a novel composition and method including a chelating agent (A) and a basic agent (B) that establishes an overall basic environment for the chelating agent (A) to effectively remove silicate-containing scale from the interior surfaces of a boiler.

DETAILED DESCRIPTION OF THE INVENTION

[0010] The composition and method of the subject invention remove scale from a substrate. More specifically, the composition and method provide for the in situ removal of scale, preferably silicate-containing scale, that is deposited on interior surfaces of various heat exchange apparatus, such as a steam boiler, during normal operation of the apparatus. To remove scale that is deposited on the interior surfaces, the substrate is exposed, in some manner, to the composition. Preferably, the composition is circulated to contact the interior surfaces. Of course, it is to be understood that the composition and method of the subject invention can be applied to any vessel in which an aqueous liquid is handled without varying the scope of the subject invention. The silicate-containing scale that is effectively removed by the composition and method disclosed herein is more

completely described as a silicate-sulfate complex of calcium, magnesium, aluminum, and other various metals. An example of such a silicate-sulfate complex is calcium silicate, CaSiO_3 .

[0011] The composition includes a chelating agent (A) and a basic agent (B). The
 5 chelating agent (A) has at least two carboxylic acid functional groups, and the basic agent (B) establishes a pH of from 7 to 14 in the composition and is selected from the group consisting of alkali metal hydroxides. The chelating agent (A), the basic agent (B), and the significance of the basic pH are described in detail below.

[0012] The chelating agent (A) is selected from the group consisting of compounds
 10 having at least two carboxylic acid functional groups (A)(i), compounds having at least two functional groups, other than carboxylic acid functional groups, that are convertible into carboxylic acid functional groups (A)(ii), and combinations thereof. It is to be understood that, without varying the scope of the subject invention, the chelating agent (A) may also be organically or inorganically-substituted. As an example, the chelating
 15 agent (A) can include halogen substituents.

[0013] The compounds having at least two carboxylic acid functional groups (A)(i) are selected from the group consisting of dicarboxylic acids, polycarboxylic acids, and combinations thereof. Various di and polycarboxylic acids include oxalic acid, adipic acid, malonic acid, succinic acid, glutaric acid, 1,2-benzenedicarboxylic acid (phthalic
 20 acid), 1,3-benzenedicarboxylic acid, 1,4-benzenedicarboxylic acid, heptanedioic acid, citric acid, and combinations thereof. In the preferred embodiment of the subject invention, the chelating agent (A) is citric acid. Other equivalent chelating agents (A) having at least two carboxylic acid functional groups include, but are not limited to, diglycolic acid, itaconic acid, malic acid, fumaric acid, glutamic acid, tartaric acid,

ethylenediaminetetraacetic acid (EDTA), nitrilotriacetic acid (NTA), ethyleneglycol-bis(beta-amino-ethyl ether)-N,N-tetraacetic acid, and combinations thereof.

[0014] The compounds having at least two functional groups, other than carboxylic acid functional groups, that are convertible into carboxylic acid functional groups (A)(ii) are selected from the group of dicarboxylic acid derivatives convertible into dicarboxylic acids by hydrolysis, polycarboxylic acid derivatives convertible into polycarboxylic acids by hydrolysis, and combinations thereof. More specifically, the dicarboxylic acid derivatives convertible into dicarboxylic acids by hydrolysis are selected from the group consisting of diacyl halides, cyclic carboxylic acid anhydrides, esters of dicarboxylic acids, dicarboxamides, and combinations thereof. As an example, phthalic anhydride, a cyclic carboxylic acid anhydride and the corresponding acid anhydride of phthalic acid, reacts with water to produce phthalic acid. Therefore, phthalic anhydride is a dicarboxylic acid derivative that is convertible into a dicarboxylic acid by hydrolysis.

[0015] The chelating agent (A) is present in the composition in an amount from 5 to 25, preferably from 10 to 20, parts by weight based on 100 parts by weight of total composition. Although not required, the chelating agent (A) is preferably in solution with water when added to the composition. A suitable solution of the chelating agent (A) in water is a 10 to 50 percent solution with water, by weight of the chelating agent (A). The most preferred solution of the chelating agent (A) in water is a 15 to 30 percent solution of the chelating agent (A) with water, by weight of the chelating agent (A).

[0016] The chelating agent (A) of the subject invention is at least a bidentate ligand. In other words, the chelating agent (A) is either a bidentate ligand or a polydentate ligand. For example, oxalic acid is a common bidentate ligand that is capable of bonding to a metal atom through two of its four oxygen atoms. Citric acid is a common polydentate

ligand that is capable of bonding to a metal atom through more than two oxygen atoms.

It is to be understood that the metal atom acted on by the bidentate or the polydentate ligand is either the calcium, magnesium, aluminum, or the other various metals contained within the silicate-containing scale that is removed by the subject composition.

- 5 [0017] As stated above, the basic agent (B) of the subject invention is an alkali metal (Group IA of the Periodic Table of Elements) hydroxide. That is, the basic agent (B) is a hydroxide of Li, Na, K, Rb, Cs, or Fr. In preferred embodiments of the subject invention, the basic agent (B) is selected from the group consisting of NaOH, KOH, and combinations thereof. In the presence of water, the basic agent (B) ionizes to produce
- 10 hydroxide ions (OH^-) which establish an overall basic environment in the composition. More specifically, the basic agent (B) establishes a pH of from 7 to 14 in the composition. Preferably, the basic agent (B) establishes a pH of from 10 to 14, more preferably from 12 to 14. The basic environment in the composition encourages precipitation of the metals contained within the silicate-containing scale as described below.
- 15 [0018] The basic agent (B) is present in an amount from 5 to 35, preferably from 10 to 25, parts by weight based on 100 parts by weight of the total composition. As with the chelating agent (A), the basic agent (B) is preferably in solution with water when added to the composition. A suitable solution of the basic agent (B) in water is a 25 to 75 percent solution with water, by weight of the basic agent (B). The most preferred
- 20 solution of the basic agent (B) in water is a 40 to 60 percent solution of the basic agent (B) with water, by weight of the basic agent (B).

[0019] To achieve the basic environment in the composition, the subject invention includes a molar excess of the basic agent (B) relative to the chelating agent (A). This molar excess is from .025 to .075, preferably from .055 to .065, excess moles of the basic

agent (B). In terms of the aqueous solutions of the chelating agent (A) and of the basic agent (B) set forth above, the volume ratio of a 20 to 25 weight percent aqueous solution of the chelating agent (A) to a 45 to 55 weight percent aqueous solution of the basic agent (B) in the composition is from 1 : 4 to 4 : 1. In the most preferred embodiment of the subject invention, the chelating agent (A) is in a 22.5 weight percent aqueous solution, the basic agent (B) is in a 50 weight percent aqueous solution, and the volume ratio of the aqueous solution of the chelating agent (A) to the aqueous solution of the basic agent (B) is 4 : 1. The volume ratio of the aqueous solutions of the chelating agent (A) and of the basic agent (B) establishes a composition having a boiling temperature of from 100 to 120, preferably from 105 to 115°C.

[0020] Due to the aqueous solutions of the chelating agent (A) and the basic agent (B), the preferred composition inherently includes water. Based on 100 parts by weight of the total composition, water is present in an amount from 50 to 75 parts by weight. In the preferred embodiment, the water is present in an amount from 55 to 70 parts by weight, and in the most preferred embodiment, the water is present in an amount from 61 to 69 parts by weight.

[0021] The composition optionally includes a monodentate ligand bonding to the metal atoms through one of its atoms. If included, the preferred monodentate ligands are selected from the group consisting of carboxylic acids having from 1 to 20 carbon atoms. Examples of suitable carboxylic acids include formic acid, acetic acid, propanoic acid, butanoic acid, pentanoic acid, benzoic acid, and equivalents thereof. One example of other suitable monodentate ligands not having carboxylic acid functionality that may be included in the composition is ammonia.

15 [0023] As described above, the basic agent (B) produces and contributes hydroxide ions in the composition. The hydroxide ions chemically attack the silicon atoms in the silicate-containing scale to disrupt the Si – O bonds. More specifically, the hydroxide ions chemically attack the silicon atoms in the silicate-sulfate complexes of the calcium, magnesium, aluminum, and other various metals to disrupt the Si – O bonds. As a result, 20 the metal atoms of the calcium, magnesium, aluminum, and other various metals that are complexed with the Si – O bonds are exposed. Next, the chelating agent (A), in the preferred embodiment citric acid, interacts with and is able to extract the exposed metal atoms from the scale. That is, the chelating agent (A), which is either a bi or polydentate ligand, bonds to the metal atoms. Finally, achieving the overall basic environment in the

composition, via the concentration of the hydroxide ions from the basic agent (B), is desirable as basic pHs enable the metal atoms to precipitate out of the composition as solid particulates of metal hydroxides. In fact, if the hydroxide ion concentration is too low, a pH of less than 7, then the effectiveness of the composition of the subject invention is diminished and it is more difficult for the metal atoms to be precipitated out of the composition. After precipitation, the solid particulates of the metal hydroxides, calcium hydroxide, magnesium hydroxide, aluminum hydroxide, and other metal hydroxides, can then be filtered from the composition. Filtering of the composition is described below.

- 10 [0024] The method disclosed in the subject invention uses the composition set forth above to remove the scale from the substrate. This method is most effective in removing silicate-containing scale previously deposited on the interior surfaces of the boiler. In this method, a first volume of the composition is circulated to contact the substrate. In terms of the steam boiler, the first volume of the composition is first charged into the boiler, and then the first volume is circulated throughout the boiler to contact the various internal surfaces of the boiler. The composition is preferably heated to from 50 to 110°C to increase effectiveness in removing the scale. More preferably, the composition is heated to from 90 to 105°C. It is also preferred that the composition is heated during circulation throughout the boiler. However, the composition may alternatively be heated by external sources prior to circulation throughout the boiler.
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[0025] The composition is circulated for from 1 to 48, preferably from 20 to 30, hours. After the appropriated circulation time, the substrate is inspected to determine if the scale has been removed. If, after inspection, it is determined that there is still additional scale deposited on the substrate that requires removal, then the first volume of the composition

can be re-circulated to remove this additional scale from the substrate. If necessary, this re-circulation step is preferably conducted for from 5 to 20, more preferably from 9 to 15, hours, to remove the additional scale. Next, the first volume of the composition is filtered to eliminate the solid particulates of the metal hydroxides. After the first volume has been filtered, the first volume is preferably supplemented with additional basic agent (B).

[0026] Alternatively, if it is determined that there is still additional scale deposited on the substrate, then a 'fresh' or second volume of the composition may be charged into the boiler and then re-circulated to remove this additional scale. As with the first volume of the composition, the second volume of the composition may be filtered to remove the solid particulates of the metal hydroxides if subsequent re-circulation is necessary.

[0027] After circulation of the composition and after an adequate inspection has determined that the scale has been removed, then certain steps are take to prepare the substrate for a return-to-service, i.e., normal operation. The substrate is rinsed with water. Next, the substrate is also passivated as is known in the art to further prepare the substrate for the return-to-service. Passivation of the substrate prevents subsequent degradation of the substrate by corrosion. A passivating solution, comprising hydrochloric acid and sodium nitrate, is circulated to contact the substrate. More specifically, the passivating solution comprises 1 weight percent HCl in water and 1 weight percent sodium nitrite, NaNO_2 , in water and is preferably circulated for from 1 to 8 hours. In the preferred embodiment where the substrate that is treated is the steam boiler, the passivating solution is circulated throughout the boiler to contact the various internal surfaces of the boiler.

[0028] The following examples illustrating the formation of the composition according to the subject invention and illustrating certain properties and results of the composition, as presented herein, are intended to illustrate and not limit the invention.

5 [0029] **Example 1**

[0030] A 25 weight percent solution of citric acid was prepared by combining 75 parts of water and 25 parts of technical grade citric acid, the chelating agent (A), and stirring until dissolved. A 50 weight percent solution of sodium hydroxide was prepared by combining 50 parts of technical grade sodium hydroxide, the basic agent (B), and 50 parts of water and stirring until dissolved. The composition was then prepared by combining 10 80 parts, by volume, of the citric acid solution and 20 parts, by volume, of the sodium hydroxide solution. The composition was then heated to 100°C, and a first volume of the composition was circulated through the boiler for twenty-four hours.

[0031] After one to two hours, the circulating composition became opaque with solid 15 particulates. These solid particulates were dissolved from the silicate-containing scale and then re-precipitated by the hydroxide ions. After twenty-four hours of circulation, the composition was removed from the boiler. The boiler was then visually inspected to determine if the scale was removed. The scale was removed. The boiler was then rinsed with fresh water. Finally, the boiler was passivated by circulating a mixture of 1 weight 20 percent hydrochloric acid and 1 weight percent sodium nitrate in the boiler for four hours.

[0032] The composition of Example 1 had a pH \geq 12, and resulted in a 10 Rating for amount of 'attack' on and removal of the silicate-containing scale previously deposited on the interior surfaces of the substrate, i.e., the boiler. Furthermore, the composition of Example 1 did not degrade the interior surfaces of the boiler during circulation.

[0033] The Ratings for the amount of attack on scale were established by repeated visual inspection of the boiler after the composition had been circulated to determine if the silicate-containing scale had been removed by reaction between the composition and the silicate-containing scale. The Ratings range from 1 to 10, with the worst rating 1
5 representing that no silicate-containing scale had been removed, and with the best rating 10 representing that all silicate-containing scale had been removed.

[0034] **Example 2**

[0035] A 25 weight percent solution of citric acid was prepared by combining 75 parts
10 of water and 25 parts of technical grade citric acid, the chelating agent (A), and stirring until dissolved. A 50 weight percent solution of potassium hydroxide was prepared by combining 50 parts of technical grade potassium hydroxide, the basic agent (B), and 50 parts of water and stirring until dissolved. The composition was then prepared by combining 80 parts, by volume, of the citric acid solution and 20 parts, by volume, of the
15 potassium hydroxide solution. The composition was then heated to 100°C, and a first volume of the composition was circulated through the boiler for twenty-four hours.

[0036] After one to two hours, the circulating composition became opaque with solid particulates. These solid particulates were dissolved from the silicate-containing scale and then re-precipitated by the hydroxide ions. After twenty-four hours of circulation,
20 the composition was removed from the boiler. The boiler was then visually inspected to determine if the scale was removed. The scale was removed. The boiler was then rinsed with fresh water. Finally, the boiler was passivated by circulating a mixture of 1 weight percent hydrochloric acid and 1 weight percent sodium nitrate in the boiler for four hours.

[0037] The composition of Example 2 had a pH of ≥ 12 , and resulted in a 10 Rating for amount of 'attack' on and removal of the silicate-containing scale previously deposited on the interior surfaces of the substrate, i.e., the boiler. Furthermore, the composition of Example 2 did not degrade the interior surfaces of the boiler during circulation.

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[0038] **Example 3**

[0039] A 20 weight percent solution of citric acid was prepared by combining 80 parts of water and 20 parts of technical grade citric acid, the chelating agent (A), and stirring until dissolved. The composition was then prepared by combining 100 parts, by volume, of the citric acid solution and 10 parts, by weight, of solid technical grade potassium hydroxide, the basic agent (B), and stirring until the potassium hydroxide was in solution. The composition was then heated to 100°C, and a first volume of the composition was circulated through the boiler for twenty-four hours.

[0040] After one to two hours, the circulating composition became opaque with solid particulates. These solid particulates were dissolved from the silicate-containing scale and then re-precipitated by the hydroxide ions. After twenty-four hours of circulation, the composition was removed from the boiler. The boiler was then visually inspected to determine if the scale was removed. The scale was removed. The boiler was then rinsed with fresh water. Finally, the boiler was passivated by circulating a mixture of 1 weight percent hydrochloric acid and 1 weight percent sodium nitrate in the boiler for four hours.

[0041] The composition of Example 3 had a pH of ≥ 12 , and resulted in a 10 Rating for amount of 'attack' on and removal of the silicate-containing scale previously deposited on the interior surfaces of the substrate, i.e., the boiler. Furthermore, the composition of Example 3 did not degrade the interior surfaces of the boiler during circulation.

[0042] Example 4

[0043] A 20 weight percent solution of citric acid was prepared by combining 80 parts of water and 20 parts of technical grade citric acid, the chelating agent (A), and stirring
5 until dissolved. The composition was then prepared by combining 100 parts, by volume, of the citric acid solution and 10 parts, by weight, of solid technical grade sodium hydroxide, the basic agent (B), and stirring until the sodium hydroxide was in solution. The composition was then heated to 100°C, and a first volume of the composition was circulated through the boiler for twenty-four hours.

10 [0044] After one to two hours, the circulating composition became opaque with solid particulates. These solid particulates were dissolved from the silicate-containing scale and then re-precipitated by the hydroxide ions. After twenty-four hours of circulation, the composition was removed from the boiler. The boiler was then visually inspected to determine if the scale was removed. The scale was removed. The boiler was then rinsed
15 with fresh water. Finally, the boiler was passivated by circulating a mixture of 1 weight percent hydrochloric acid and 1 weight percent sodium nitrate in the boiler for four hours.

[0045] The composition of Example 4 had a pH of ≥ 12 , and resulted in a 10 Rating for amount of 'attack' on and removal of the silicate-containing scale previously deposited on the interior surfaces of the substrate, i.e., the boiler. Furthermore, the composition of
20 Example 4 did not degrade the interior surfaces of the boiler during circulation.

[0046] Example 5

[0047] A 20 weight percent solution of citric acid was prepared by combining 80 parts of water and 20 parts of technical grade citric acid, the chelating agent (A), and stirring

until dissolved. The composition was then prepared by combining 100 parts, by volume, of the citric acid solution and 20 parts, by weight, of solid technical grade sodium hydroxide, the basic agent (B), and stirring until the sodium hydroxide was in solution. The composition was then heated to 100°C, and a first volume of the composition was
5 circulated through the boiler for twenty-four hours.

[0048] After one to two hours, the circulating composition became opaque with solid particulates. These solid particulates were dissolved from the silicate-containing scale and then re-precipitated by the hydroxide ions. After twenty-four hours of circulation, the composition was removed from the boiler. The boiler was then visually inspected to
10 determine if the scale was removed. The scale was almost all removed. However, it was determined that additional circulation was needed. As such, a second volume or 'fresh' volume of the composition was heated to 100°C, and the second volume of the composition was circulated through the boiler for an additional twelve hours. The boiler was then visually inspected again to if the scale was removed. The scale was all
15 removed. The boiler was then rinsed with fresh water. Finally, the boiler was passivated by circulating a mixture of 1 weight percent hydrochloric acid and 1 weight percent sodium nitrate in the boiler for four hours.

[0049] The composition of Example 5 had a pH of ≥ 12 , and resulted in a 10 Rating for amount of 'attack' on and removal of the silicate-containing scale previously deposited
20 on the interior surfaces of the substrate, i.e., the boiler. Furthermore, the composition of Example 5 did not degrade the interior surfaces of the boiler during circulation.

[0050] Example 6

[0051] A 20 weight percent solution of citric acid was prepared by combining 80 parts of water and 20 parts of technical grade citric acid, the chelating agent (A), and stirring until dissolved. The composition was then prepared by combining 100 parts, by volume, of the citric acid solution and 10 parts, by weight, of solid technical grade sodium hydroxide, the basic agent (B), and stirring until the sodium hydroxide was in solution. The composition was then heated to 100°C and circulated through the boiler for twenty-four hours.

[0052] After one to two hours, the circulating composition became opaque with solid particulates. These solid particulates were dissolved from the silicate-containing scale and then re-precipitated by the hydroxide ions. After twenty-four hours of circulation, the composition was removed from the boiler. The boiler was then visually inspected to determine if the scale was removed. The scale was almost all removed. However, it was determined that additional circulation was needed. As such, the first volume of the composition was filtered to eliminate any solid particulates in the first volume. After filtering, the first volume of the composition was supplemented with an additional 10 parts, by weight, of the solid technical grade sodium hydroxide to ensure the basic environment. After supplementing, the 'refreshed' first volume was heated to 100°C and then was re-circulated through the boiler for an additional twelve hours. The boiler was then visually inspected again to if the scale was removed. The scale was all removed. The boiler was then rinsed with fresh water. Finally, the boiler was passivated by circulating a mixture of 1 weight percent hydrochloric acid and 1 weight percent sodium nitrate in the boiler for four hours.

[0053] The composition of Example 6 had a pH of ≥ 12 , and resulted in a 10 Rating for amount of 'attack' on and removal of the silicate-containing scale previously deposited on the interior surfaces of the substrate, i.e., the boiler. Furthermore, the composition of Example 6 did not degrade the interior surfaces of the boiler during circulation.

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[0054] Comparative Examples

[0055] The several comparative examples detailed in the following table were prepared by adding the composition parts by weight, unless otherwise indicated. Only the non-aqueous parts of the comparative examples are detailed. That is, in all cases, the remainder of the composition for the comparative examples is water, by weight. Examples 7 and 8 (listed in the table as Comparative Examples 19 and 20, respectively), detailing compositions according to the subject invention, are included in the table for comparative purposes. Examples 7 and 8 were prepared as described above in Example 1.

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<u>Comp. Example No.</u>	<u>Composition</u>	<u>Result</u>
1	32% hydrochloric acid	No attack on silicate-containing scale. Degrades interior surfaces of boiler. Rating = 1
2	100% acetic acid	No attack on silicate-containing scale. Degrades interior surfaces of boiler. Rating = 1
3	35% acetic acid	No attack on silicate-containing scale. Degrades interior surfaces of boiler. Rating = 1
4	98% sulfuric acid	No attack on silicate-containing scale. Degrades interior surfaces of boiler. Rating = 1
5	33% sulfuric acid	No attack on silicate-containing scale. Degrades interior surfaces of boiler. Rating = 1
6	85% phosphoric acid	No attack on silicate-containing scale. Degrades interior surfaces of boiler.

		Rating = 1
7	35% phosphoric acid	No attack on silicate-containing scale. Degrades interior surfaces of boiler. Rating = 1
8	70% nitric acid	No attack on silicate-containing scale. Degrades interior surfaces of boiler. Rating = 1
9	35% nitric acid	No attack on silicate-containing scale. Degrades interior surfaces of boiler. Rating = 1
10	35% nitric acid / 16% hydrochloric acid	No attack on silicate-containing scale. Degrades interior surfaces of boiler. Rating = 1
11	10% hydrofluoric acid	No attack on silicate-containing scale. Degrades interior surfaces of boiler. Rating = 1
12	32% hydrochloric acid / 10% ammonium fluoride	No attack on silicate-containing scale. Degrades interior surfaces of boiler. Rating = 1
13	10% acetic acid / 8% EDTA	No reaction. Rating = 1
14	20% acetic acid / 20% citric acid	No reaction. Rating = 1
15	5% sodium silicate	No reaction. Rating = 1
16	10% EDTA	No reaction. Rating = 1
17	10% sodium hydroxide / 8% sodium borate	No reaction. Rating = 1
18	50% sodium hydroxide	Minor attack, i.e., reaction, on silicate-containing scale. Rating = 3
19	<u>Example 7</u> 10% sodium hydroxide / 8% EDTA	Almost all silicate-containing scale removed. Rating = 8
20	<u>Example 8</u> 20% sodium hydroxide / 20% citric acid	All silicate-containing scale removed. Rating = 10

[0056] The results included in the above table indicate that Comparative Examples 1 – 18 did not successfully remove the silicate-containing scale from the substrate. In fact, in most instances, the comparative example compositions actually degraded the interior surfaces of the boiler substrate. On the other hand, Comparative Example 19, having a Rating of 8, removed almost all silicate-containing scale, and Comparative Example 20, having a Rating 10, removed all of the silicate-containing scale from the substrate.

[0057] The invention has been described in an illustrative manner, and it is to be

1. The first part of the document is a list of names and their corresponding addresses. The names are: "John Doe", "Jane Smith", "Bob Johnson", "Alice Brown", "Charlie White", "David Green", "Eve Black", "Frank Gray", "Grace Hall", "Henry King", "Ivy Lee", "Jack Miller", "Karen Wilson", "Leo Young", "Mia Adams", "Noah Baker", "Olivia Carter", "Peter Davis", "Quinn Evans", "Ruth Foster", "Samuel Gibson", "Tina Harris", "Uma Ives", "Victor Jones", "Wendy Kelly", "Xavier Lewis", "Yara Martin", "Zoe Nelson". The addresses are: "123 Main St, New York, NY 10001", "456 Elm St, Los Angeles, CA 90001", "789 Oak St, Chicago, IL 60601", "101 Pine St, San Francisco, CA 94101", "202 Cedar St, Boston, MA 02101", "303 Birch St, Philadelphia, PA 19101", "404 Spruce St, Washington, DC 20001", "505 Willow St, Houston, TX 77001", "606 Ash St, Phoenix, AZ 85001", "707 Hickory St, San Antonio, TX 78201", "808 Cypress St, Fort Worth, TX 76101", "909 Redwood St, Dallas, TX 75201", "1010 Sycamore St, Austin, TX 78701", "1111 Magnolia St, Jacksonville, FL 32201", "1212 Dogwood St, Orlando, FL 32801", "1313 Palm St, Miami, FL 33101", "1414 Peach St, Tampa, FL 33601", "1515 Apple St, St. Petersburg, FL 33701", "1616 Cherry St, Clearwater, FL 34601", "1717 Plum St, Lakeland, FL 33801", "1818 Pear St, Palmdale, CA 93550", "1919 Olive St, Lancaster, CA 93534", "2020 Lemon St, Hanford, CA 93231", "2121 Lime St, Clovis, CA 93611", "2222 Grape St, Merced, CA 95340", "2323 Strawberry St, Modesto, CA 95201", "2424 Blueberry St, Yuba City, TX 75901", "2525 Raspberry St, Baytown, TX 77601", "2626 Blackberry St, Houston, TX 77001", "2727 Elderberry St, The Woodlands, TX 77380", "2828 Currant St, Springdale, TX 75782", "2929 Fig St, Rockwall, TX 75087", "3030 Mulberry St, Irving, TX 75039", "3131 Elderberry St, Carrollton, TX 75006", "3232 Raspberry St, Frisco, TX 75034", "3333 Strawberry St, McKinney, TX 75069", "3434 Blueberry St, Denton, TX 76201", "3535 Blackberry St, Carrollton, TX 75006", "3636 Elderberry St, Lewisville, TX 75057", "3737 Currant St, Flower Mound, TX 76041", "3838 Fig St, Georgetown, TX 78626", "3939 Mulberry St, Round Rock, TX 78664", "4040 Lemon St, Austin, TX 78701", "4141 Lime St, Cedar Park, TX 78613", "4242 Grape St, Leander, TX 78641", "4343 Strawberry St, Killebrew, TX 78949", "4444 Blueberry St, Temple, TX 76788", "4545 Blackberry St, Belton, TX 77705", "4646 Elderberry St, San Marcos, TX 78681", "4747 Currant St, New Braunfels, TX 78130", "4848 Fig St, Seguin, TX 78155", "4949 Mulberry St, Comstock, TX 77956", "5050 Lemon St, Del Rio, TX 78840", "5151 Lime St, El Paso, TX 79901", "5252 Grape St, El Paso, TX 79901", "5353 Strawberry St, El Paso, TX 79901", "5454 Blueberry St, El Paso, TX 79901", "5555 Blackberry St, El Paso, TX 79901", "5656 Elderberry St, El Paso, TX 79901", "5757 Currant St, El Paso, TX 79901", "5858 Fig St, El Paso, TX 79901", "5959 Mulberry St, El Paso, TX 79901", "6060 Lemon St, El Paso, TX 79901".